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The first law of thermodynamics is a version of the law of conservation of energy, adapted for thermodynamic processes, distinguishing two kinds of transfer of energy, as heat and as thermodynamic work, and relating them to a function of a body's state, called Internal energy.. The law of conservation of energy states that the total energy of an isolated system is constant; energy can be ...

First law of thermodynamics - Wikipedia

Examples of thermal efficiencies are 4% for steam locomotives, 18%-30% for internal combustion engines, and 30% for modern power stations. From these examples it can be seen that the conversion of heat into work involves considerable losses in the available energy, and indeed, the efficiencies of most heat engines are generally much less than 100%.

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